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About the origin of the large Stokes shift in aminoalkyl substituted heptamethine cyanine dyes†

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Aminoalkyl-substituted heptamethine cyanine dyes are characterized by a large Stokes shift, an uncommon feature for cyanine molecules yet very promising for their application as fluorescent probes in bioimaging and as light harvesting antennas in biohybrid systems for solar energy conversion. The origin of this photophysical feature has not been adequately explored so far, and a combined experimental and theoretical work is herein provided to shed light on the role played by the central aminoalkyl substituent bound to the heptamethine cyanine backbone in defining the unusual properties of the dye. The spectra recorded in solvents of different polarities point to a marginal role of the medium in the definition of the Stokes shift, which conversely can be ascribed to the relaxation of the molecular geometry upon photoexcitation. This hypothesis is supported by an extensive theoretical investigation of the ground and excited states of the dye. TD-DFT results on the aminoalkyl-substituted dye and its unsubstituted precursor demonstrate a very similar cyanine-like structure for both molecules in the relaxed excited state. Conversely, in the ground state the amino substitution disrupts the conjugation in the polymethine chain, leading to a broken-symmetry, non-planar structure.

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Introduction

Cyanines are a popular class of synthetic π -conjugated dyes characterized by a symmetric and charged structure with a polymethine chain bound to two peripheral electron donor (D) or acceptor (A) groups. The scientific interest towards cyanine dyes is mainly related to their spectral properties, such as high molar extinction coefficients and good fluorescence quantum yields, which make them excellent fluorescent probes for biological applications.^{1,2} The properties of cyanine dyes can be easily tuned by appropriately tailoring their chemical structure and functionalization by means of organic synthesis. The length of the polymethine chain and the nature of the peripheral groups can be modified to change the spectral position of absorption and emission bands. Furthermore, substituents can be designed to improve the solubility and/or provide chemical affinity to specific biotargets.^{3,4} Cyanines with short polymethine chains

typically show negligible solvatochromism, narrow absorption and emission bands with weak vibronic structures and very small Stokes shifts.⁵ On the opposite, cyanines with long polymethine chains are prone to symmetry breaking in the ground state, a phenomenon responsible for their unusual spectral properties:^{6–9} the position and the shape of absorption bands of long cyanines become strongly solvent-dependent, with broad and blue-shifted bands in polar solvents, while fluorescence bands maintain the weak solvatochromism and the vibronic structure typical of short cyanines. Essential-state models rationalize the spectral behaviour of both symmetry-preserving and symmetry-breaking cyanine dyes in a single theoretical framework.^{8,10,11}

Heptamethine cyanines are promising for bioapplications; their spectra are in fact located in the near-infrared region, where the ability to penetrate biological tissues is the highest and the autofluorescence is marginal. Unfortunately, two main drawbacks hinder their use as fluorescent bioprobes: poor photostability and a very small Stokes shift (about 0.05 eV), the latter causing fluorescence self-absorption and quenching. Photostability is improved by introducing a chlorocyclohexenyl ring in the central position of the polymethine chain,¹² and this also enhances the fluorescence quantum yield and provides a reactive site. This reactive site has been exploited to introduce a series of substituents on the central cyclohexenyl ring via nucleophilic substitution of the chlorine atom with amines,¹³

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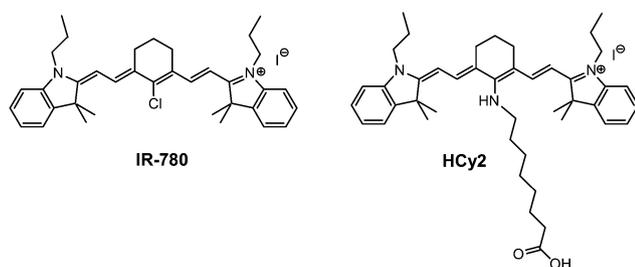


thiols,^{14–16} alcohols^{17,18} and phenyls.^{19,20} In particular, amino-substituted heptamethine cyanine dyes obtained by substitution of the chlorine atom show a huge Stokes shift (larger than 0.3 eV),¹³ maintaining or even improving the fluorescence properties typical of heptamethine dyes. These molecules have been exploited as fluorescent probes for bioimaging^{21,22} and in artificial light harvesting antennas for solar energy conversion.^{23–25}

The mechanism at the basis of the spectroscopic effects observed upon substitution of the chlorine atom with an amino-alkyl group has not been fully addressed in the literature. Different research groups proposed an interplay between localized excitation and intermolecular charge transfer, but clear experimental evidence for this interplay is missing.^{13,23} The ion-pairing effect was also invoked for heptamethine cyanines showing different behaviours according to the counterion and/or the polarity of the solvent.^{19,26} The nitrogen group in the central core definitely affects the ground-state electronic configuration, favouring the presence of the charge on the central position.^{27,28} However, a thorough investigation of the first excited state and of its relaxation after vertical excitation is necessary to understand the origin of the large Stokes shift.

In this paper, a comparison between two cyanine dyes with different spectral properties is presented (Scheme 1): **IR-780** is a commercially available dye bearing a chlorocyclohexenyl ring in the central position of the heptamethine cyanine structure, showing negligible (≈ 0.04 eV) Stokes shift; conversely **HCy2**, the amino-alkyl functionalized analogue obtained by nucleophilic substitution of the chlorine atom in **IR-780** shows a large (≈ 0.3 eV) Stokes shift. **HCy2** has been selected and synthesized in the frame of studies on biohybrid systems for solar energy conversion based on the covalent bioconjugation of organic light harvesting antennas to photoenzymes.^{23–25} The specific octyl chain of **HCy2** was chosen as a suitable spacer between the chromophore and the protein, in order to prevent the sterical hindrance of **HCy2** on the photoenzyme surface, thus preserving the enzymatic photoactivity.^{29,30}

A combined experimental and theoretical work sheds light on the origin of this large Stokes shift: the nitrogen atom linked to the cyclohexenyl ring strongly affects the conjugation of the polymethine chain, leading to a broken-symmetry ground state. On the opposite, the symmetry is partially restored in the first excited state (upon relaxation), leading to a large structural re-arrangement of the molecule that explains the observed large Stokes shift.



Scheme 1 Molecular structures of **IR-780** and **HCy2**.

Methods

Synthesis

As reported in previous work,²⁵ **HCy2** was synthesized by reacting overnight **IR-780** with 8-aminooctanoic acid in anhydrous *N,N*-dimethylformamide (DMF), under a nitrogen atmosphere at 80 °C. DMF was then removed by dissolving the reaction mixture in dichloromethane and washing it with water. After drying the organic phase with anhydrous Na₂SO₄, dichloromethane was distilled at a reduced pressure and **HCy2** was isolated as a blue solid in 48% yield by reverse phase C-18 column chromatography of the crude product using a mixture of water and acetonitrile (6 : 4 v/v) as the mobile phase.

Spectroscopic measurements

Spectra grade or HPLC solvents were used to prepare solutions for absorption and emission measurements. Absorption spectra were recorded with a PerkinElmer Lambda650 spectrophotometer, while emission spectra were recorded with a FLS1000 Edinburgh Fluorometer equipped with a R5509-72 (Hamamatsu) NIR-PMT for detection in the near infrared region. Fluorescence spectra were recorded on dilute samples (with absorbance lower than 0.1). Fluorescence anisotropy spectra were recorded using a Horiba Jobin Ivon Fluoromax-3 instrument, equipped with polarizers. The sample was cooled using the Horiba FL-1013 Liquid Nitrogen Dewar assembly.

Quantum chemical calculations

All the reported quantum chemical calculations were performed using the Gaussian16 software.³¹ Ground-state geometry optimization was obtained at the DFT level of theory using ω B97xD, B3LYP and M062X functionals and the 6-31g(d) basis set in toluene, DCM and DMSO, adopting the standard PCM model for solvation. Frequencies were computed after optimization using the same method. Vertical energies were calculated at the TD-DFT level of theory, adopting the same functionals and basis set. Solvation was accounted for in the framework of the corrected linear response theory.^{32,33} The relaxation of the first excited state towards its minimum energy was computed at the TD-DFT level, using the same functionals and basis set adopted for the ground state. Emission transition energies were calculated as vertical transitions from the relaxed vertical excited state (accounting for the corrected linear response theory for solvation).

Results and discussion

Spectroscopic measurements

The absorption and emission spectra recorded in solvents of different polarities (DCM and DMSO) are shown in Fig. 1 and the main data are summarized in Table 1. Absorption and fluorescence spectra of **IR-780** are marginally affected by the solvent: the main absorption band is observed at 790 (796) nm in DCM (DMSO), and the Stokes shift is very small, with the emission maximum detected at 808 (816) nm. Absorption and emission bands are narrow, with a weak vibronic shoulder in both solvents. The main absorption band of **HCy2** at 665 (636) nm in



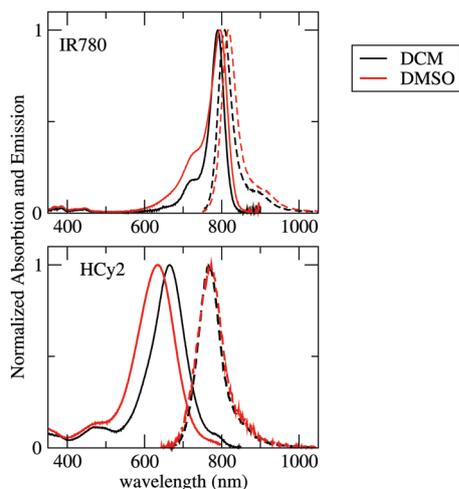


Fig. 1 Absorption (continuous lines) and emission (dashed lines) spectra of **IR-780** and **HCy2** in DCM (black lines) and DMSO (red lines).

Table 1 Experimental spectroscopic properties

Molecule	Solvent	$\lambda_{\text{max}}^{\text{abs}}$ [nm]	$\lambda_{\text{max}}^{\text{em}}$ [nm]	Stokes shift [eV]
IR-780	DCM	790	808	0.03
	DMSO	796	816	0.04
HCy2	DCM	665	768	0.25
	DMSO	636	773	0.35

DCM (DMSO) is considerably blue-shifted with respect to that of **IR-780**, while the emission at 768 (773) nm is only slightly blue shifted with respect to that of **IR-780**. **HCy2** shows a large Stokes shift, which increases with the solvent polarity from 0.25 eV in DCM to 0.35 eV in DMSO. **HCy2** absorption bandshapes are broad, with an unresolved vibronic structure, while emission has a narrow bandshape, similar to that of **IR-780**. Neither **IR-780** nor **HCy2** are soluble in non-polar solvents.

Fluorescence anisotropy spectra provide details about the relative orientation of transition dipole moments of absorption and emission processes and therefore can be exploited to assess the possible contribution of different electronic excitations to the same absorption band, useful information when dealing with broad absorption spectra as for **HCy2**. The fluorescence excitation anisotropy spectrum recorded for **HCy2** in glassy ethanol (EtOH) at 77 K is shown in Fig. 2. The fluorescence anisotropy signal is flat within the lowest excitation band, suggesting that a single electronic transition is responsible for the absorption band. The measured anisotropy is close to 0.3, meaning that transition dipole moments of emission and absorption form an angle close to 24° . The Stokes shift measured at 77 K is somewhat reduced compared to that at room temperature (Fig. 2), but is still sizeable. This residual Stokes shift measured in the glassy solvent can be safely attributed to the relaxation of the molecular geometry after photoexcitation.

Quantum chemical calculations

DFT/TD-DFT calculations (see the Methods section for technical details) were performed to investigate the different spectroscopic

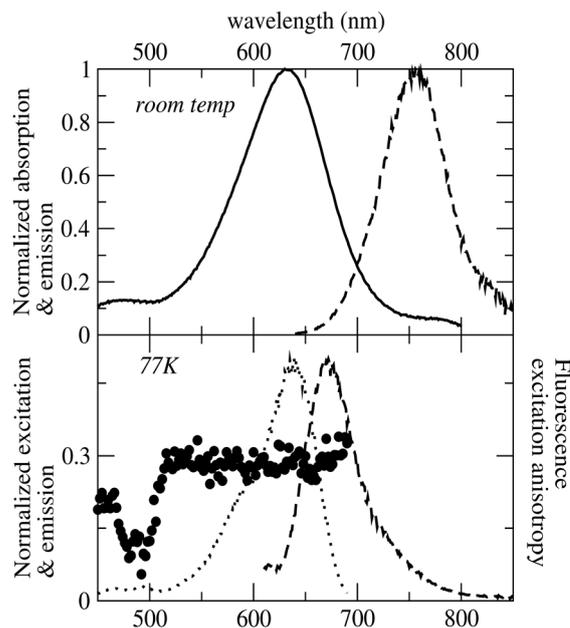


Fig. 2 Top panel: Absorption and emission of **HCy2** at room temperature in EtOH. Bottom panel: Excitation, emission and fluorescence excitation anisotropy spectra of **HCy2** in EtOH at 77 K.

behaviours of **IR-780** and **HCy2**, and in particular to understand the origin of the large Stokes shift in **HCy2**. The ground state structure of both **IR-780** and **HCy2** was optimized in DCM and DMSO (adopting the PCM scheme at the DFT level of theory). Vertical excitations were calculated by TD-DFT adopting the corrected Linear Response theory for the solvent.³² Optimization of the first excited state was performed at the same level of theory. In this section, the results obtained with ω B97xD functional are discussed. Different functionals (B3LYP and M062X) were tested, providing qualitatively similar results (see the ESI[†]), confirming the robustness of the approach. To speed up the calculation, without affecting the results, computations were performed on simplified molecular structures (Fig. 3), where the long alkyl chains are substituted by shorter groups. The carboxylic group in **HCy2**, (see Scheme 1) marginally affects its structural and optical properties, as discussed in the ESI[†].

The selected ground state properties of **IR-780** and **HCy2** are summarised in Table 2 (the dihedral angles δ_1 and δ_2 and the bond length alternations, BLA_1 and BLA_2 are defined in Fig. 3). The dihedral angles δ_1 and δ_2 provide information about the planarity of the molecule. The ground state of **IR-780** has a planar structure, with δ_1 and δ_2 close to 180° . On the opposite, **HCy2** is distinctly non-planar in the ground state, with both δ_1 and δ_2 appreciably different from 180° . Bond length alternations, $BLA_{1/2}$ are also calculated: a small BLA implies a good electronic delocalization along the chain (strong cyanine character), while a large BLA implies alternating single and double bonds. Moreover, dissimilar BLAs in the two molecular arms clearly point to symmetry breaking. In **IR-780**, BLA_1 and BLA_2 are both close to 0, meaning that all bonds in the polymethine chain have a very similar character, intermediate between single and double bonds. BLA_1 and BLA_2 are larger in **HCy2**, with the bonds relevant



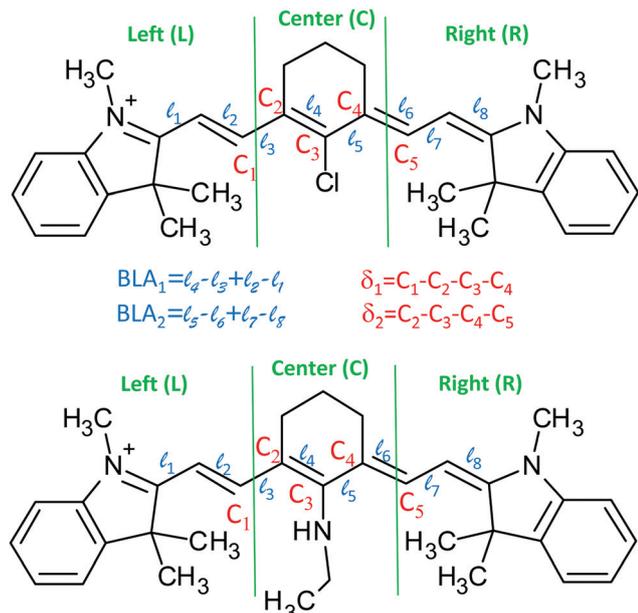


Fig. 3 Simplified molecular structures of **IR-780** (top) and **HCy2** (bottom) adopted for quantum chemical calculations. The dihedral angles, δ_1 and δ_2 , and the bond length alternations, BLA_1 and BLA_2 , are also defined. The green vertical lines mark the fragmentation of the molecules into a left and right arm and a central region.

Table 2 Selected ground state (GS) and the first excited state (1EX) properties of **IR-780** and **HCy2** in two different solvents

Molecule	Solvent	δ_1 (°)	δ_2 (°)	BLA_1 (Å)	BLA_2 (Å)	l_4 (Å)	l_5 (Å)
IR-780 (GS)	DCM	174	-174	0.015	0.015	1.408	1.408
	DMSO	174	-174	0.015	0.015	1.407	1.407
HCy2 (GS)	DCM	-162	-154	0.110	0.086	1.444	1.434
	DMSO	-162	-154	0.113	0.089	1.445	1.436
IR-780 (1EX)	DCM	174	-174	0.013	0.013	1.414	1.414
	DMSO	174	-174	0.011	0.011	1.414	1.414
HCy2 (1EX)	DCM	-173	-170	0.025	0.024	1.420	1.422
	DMSO	-174	-171	0.023	0.022	1.419	1.422

to l_1, l_3, l_6 and l_8 having a predominant character of double bonds, and l_2, l_4, l_5 and l_7 having a predominant single-bond character. Furthermore, BLA_1 differs from BLA_2 , pointing to a broken-symmetry ground state. The asymmetry is also confirmed by the length of the two central bonds of the polymethine chain (l_4 and l_5) that in **HCy2** are non-equivalent and significantly longer than in **IR-780**.

To discuss the charge distribution in the molecules, we have arbitrarily divided each molecule into three fragments (left, center and right), as defined in Fig. 3. The cumulative charges, calculated for the three fragments as the sum of Hirshfeld atomic charges, are given in Table 3. For **IR-780**, the charge is symmetrically distributed over the two molecular arms, and is mainly localized on the peripheral groups of the molecule. The structural asymmetry of **HCy2** in the ground state shows up with a non-symmetric charge distribution in the two arms. Moreover, in **HCy2** the charge residing in the central fragment is comparable to that in the peripheral subunits, in agreement with previous results by Pascal *et al.*²⁷ The dyes are not soluble

Table 3 Ground-state and the first excited state (vertical and relaxed) charge distributions. The three subunits left, central and right are defined in (Fig. 3)

Molecule	Solvent	Left	Central	Right
IR-780 (GS)	DCM	0.476	0.048	0.476
	DMSO	0.464	0.072	0.464
HCy2 (GS)	DCM	0.360	0.250	0.390
	DMSO	0.322	0.322	0.356
IR-780 (1EX-Vertical)	DCM	0.465	0.070	0.465
	DMSO	0.439	0.122	0.439
HCy2 (1EX-Vertical)	DCM	0.405	0.188	0.407
	DMSO	0.372	0.254	0.373
IR-780 (1EX-Relaxed)	DCM	0.464	0.072	0.464
	DMSO	0.441	0.118	0.441
HCy2 (1EX-Relaxed)	DCM	0.430	0.139	0.431
	DMSO	0.410	0.188	0.402

in non-polar solvents, but calculations run in toluene (see the ESI^\dagger) confirm that the solvent polarity marginally affects the ground state properties of either dye.

Table 4 summarizes the TD-DFT results for the lowest energy vertical excitation. The second electronic transition is calculated at much larger energies for both molecules, with an energy gap between the first and the second transition larger than 1.5 eV for **IR-780** and 0.9 eV for **HCy2**. Therefore, the second electronic transition is irrelevant to the present discussion. The calculated transition energies are blue-shifted compared to the experiment, a well-known issue for cyanine dyes.³⁴ The TD-DFT calculations in toluene (see the ESI^\dagger) demonstrate negligible effects of solvent polarity on the absorption spectra of **IR-780**. The calculated absorption spectra in **HCy2** slightly blue shift as the solvent polarity increases, in agreement with the experimental data. The lowest transition is well described as a HOMO–LUMO transition, and is delocalized over the whole molecule, as shown in Fig. 4. The HOMO and LUMO of **IR-780** are perfectly symmetric, while the symmetry is lost in the orbitals of **HCy2**. The decrease of symmetry has important consequences in the energy positioning of the orbitals: the energy gap between the HOMO and LUMO is increased in **HCy2**, leading to a blue shift of the absorption band.

The geometry of the first excited state was optimized using the same level of theory adopted for the ground state and accounting for the solvent in the corrected linear response theory of PCM. The solvent polarity barely affects the emission energy of either **IR-780** or **HCy2** (Table 4). The main qualitative difference between the two molecules concerns the Stokes shift, which, in excellent agreement with the experiment, is predicted to be very small for **IR-780** (less than 0.1 eV) and becomes large (more than 0.3 eV) in **HCy2**.

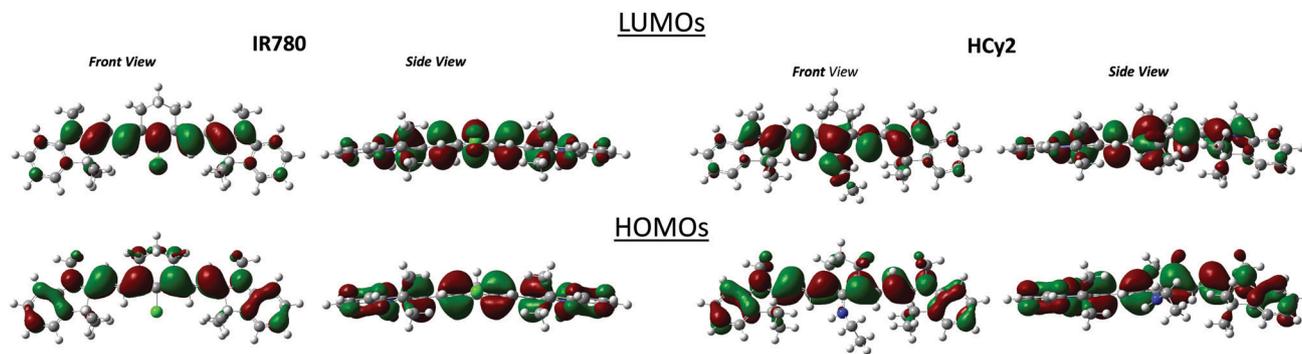
We are in the position to clarify the origin of the large Stokes shift observed for **HCy2**. The role of solvent polarity is marginal, since comparable Stokes shifts are calculated in DCM and DMSO (see Table 4) as well as in toluene (see the ESI^\dagger). The origin of Stokes shift is therefore safely ascribed to the geometry relaxation in the excited state.

Tables 2 and 3 summarize the computational results about the geometry and charge distribution in the vertical and relaxed excited states of the two molecules. Important structural



Table 4 Calculated spectroscopic properties

Molecule	Solvent	Absorption [nm/eV]	Emission [nm/eV]	Stokes shift [eV]
IR-780	DCM	556/2.23 (H–L 0.69; H–1 → L+1 –0.13)	574/2.16 (H–L 0.69; H–1 → L+1 0.13)	0.07
	DMSO	556/2.23 (H–L 0.69; H–1 → L+1 0.13)	577/2.15 (H–L 0.69; H–1 → L+1 –0.12)	0.08
HCy2	DCM	482/2.57 (H–L 0.67; H–1 → L+1 –0.16)	554/2.24 (H–L 0.69; H–1 → L+1 0.13)	0.33
	DMSO	481/2.58 (H–L 0.67; H–1 → L+1 –0.16)	556/2.23 (H–L 0.69; H–1 → L+1 0.13)	0.35

Fig. 4 The HOMO and LUMO of **IR-780** (left) and **HCy2** (right) in DMSO (equilibrium ground-state geometry).

changes upon excitation are observed for **HCy2**: (1) the dihedral angles δ_1 and δ_2 increase upon excitation and, becoming close to 180° , point to planarization of the molecule in the excited state; (2) both BLA_1 and BLA_2 significantly reduce and become comparable; and (3) the bond-lengths ℓ_4 and ℓ_5 also reduce and become comparable. Overall, the first excited state of **HCy2** has a nearly planar and symmetric structure with an almost equalized bond-length, as a typical cyanine-like system. These structural changes in the excited state with respect to the ground state are responsible for the large observed Stokes shift in **HCy2**. On the opposite, no significant structural changes are observed in the first excited state of **IR-780** with respect to the ground state, reflecting a small Stokes shift.

In **HCy2**, the charge distribution also undergoes significant changes in the excited state, as reported in Table 3. After excitation (the vertical excited state), the charge redistributes almost symmetrically over the two arms, and a significant charge depletion is observed in the central core of the molecule with respect to the ground state. The charge depletion in the central core of the molecule further increases after solvent relaxation, and in the relaxed excited state, charge distributions of **HCy2** and **IR-780** are very similar.

Quite interestingly, while important spectroscopic effects are observed upon substituting the central chlorine with amino-groups, the Stokes shift remains small in the case of substitution with thiols or alcohols.^{14,16–18} These experimental observations suggest that only the nitrogen atom has the ability to conjugate with the polymethine chain. From DFT calculations, in fact, the central C–N bond shows an intermediate character between the single and double bond (the C–N bond length in the ground state of the modified **HCy2** is ≈ 1.35 Å). On the opposite, the DFT calculations on analogous structures with $-\text{OCH}_3$ and $-\text{SCH}_3$ substituents predict a predominant character of the single bond (see the ESI[†]).

Conclusions

The chemical structures of **HCy2** and **IR-780** heptamethine cyanine-dyes only differ in the nature of the unit bound to the central carbon atom of the polymethine chain: **IR-780** bears a chlorine atom, whereas **HCy2** an amino alkyl chain. This single substitution strongly affects the optical properties of the dyes, in particular their Stokes shift.

The absorption spectra of **HCy2** are distinctly different from those of typical cyanine-dyes like **IR-780**. The absorption spectrum is strongly blue-shifted (more than 0.3 eV) and much broader compared to that of **IR-780**. On the opposite, the emission of **HCy2** has a similar bandshape with respect to **IR-780**, and is only marginally blue-shifted (around 0.07 eV), resulting in a very large Stokes shift for **HCy2**. Long polymethine dyes are prone to symmetry breaking in the ground state:⁸ broken-symmetry cyanines show very broad absorption bands and a large Stokes shift, spectral features that are also recognized in **HCy2**. However, symmetry breaking in long polymethine dyes is typically induced by polar solvation, while absorption and fluorescence spectra of **HCy2** are marginally affected by the solvent.

Quantum chemical calculations suggest that in the ground state of **HCy2**, the central C–N bond has an intermediate character between the single and double bonds. The presence of such a strong bond affects the geometry and conjugation of the polymethine chain: the ground state structure of **HCy2** is non-planar and non-symmetric. In the ground state, the charge is preferentially localized in the central part of the molecule,²⁷ and the charge distribution is not symmetric between the two molecular arms. Overall, looking at the ground state properties, **HCy2** cannot be described as a cyanine-dye, but upon excitation, the molecule becomes a typical cyanine-dye: the relaxed excited state has an almost planar and symmetric structure, whose quasi-symmetric charge distribution resembles the charge distribution in **IR-780**.



The origin of the Stokes shift of amino-substituted polymethine dyes is ascribed to the important structural changes occurring after photoexcitation. Symmetry breaking plays a key role in the observation of the large Stokes shift in symmetric structures: as experimentally observed and rationalized in terms of essential-state models,^{8,11} large Stokes shifts (strongly dependent on solvent polarity) are observed when symmetry breaking occurs in the first excited state or in the ground state.^{8,10} Here a different type of ground-state symmetry breaking is observed that originates from a structural distortion due to the strong conjugation of the amino group with the polymethine chain: this structural distortion does not depend on solvent polarity, and is almost lost in the first excited state, giving broad and blue-shifted absorption spectra, and cyanine-like narrow emission spectra.

Conflicts of interest

There are no conflicts to declare.

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